

857.

Problem 43.79P (HRW)

We have to calculate (a) the energy needed to remove a proton from a ^{121}Sb nucleus, and (b) the energy needed to remove a proton from the resulting ^{120}Sn nucleus. The needed atomic masses are

^{121}Sb	120.9038 u
^{120}Sn	119.9022 u
^{119}In	118.9058 u.



Solution:

(a)

We have to calculate the energy needed to remove a proton from a ^{121}Sb nucleus. We note that the atomic number of antimony ^{121}Sb is 51 and therefore an antimony atom has 51 electrons.

The atomic number of ^{120}Sn is 50 and so a ^{120}Sn atom contains 50 electrons.

Therefore, the energy required for removing a proton from a ^{121}Sb nucleus will be

$$\begin{aligned}
-Q &= \left(\left(m(^{120}\text{Sn}) - 50m_e \right) + m(p) - \left(m(^{121}\text{Sb}) - 51m_e \right) \right) c^2 \\
&= \left(m(^{120}\text{Sn}) + m(^1\text{H}) - m(^{121}\text{Sb}) \right) c^2 \\
&= (119.9022 + 1.007825 - 120.9038) \text{uc}^2 \\
&= 0.006225 \text{uc}^2 = 0.006225 \text{uc}^2 \times 931.5 \text{ MeV} \\
&= 5.798 \text{ MeV}.
\end{aligned}$$

(b)

We have to calculate the energy needed to remove a proton from a ^{120}Sn nucleus. We note that the atomic number of indium ^{119}In is 49 and so an ^{119}In atom contains 49 electrons.

Therefore, the energy required for removing a proton from a ^{120}Sn nucleus will be

$$\begin{aligned}
-Q &= \left(\left(m(^{119}\text{In}) - 49m_e \right) + m(p) - \left(m(^{120}\text{Sn}) - 50m_e \right) \right) c^2 \\
&= \left(m(^{119}\text{In}) + m(^1\text{H}) - m(^{120}\text{Sn}) \right) c^2 \\
&= (118.9058 + 1.007825 - 119.9022) \text{uc}^2 \\
&= 0.011425 \text{uc}^2 = 0.011425 \text{uc}^2 \times 931.5 \text{ MeV} \\
&= 10.639 \text{ MeV}.
\end{aligned}$$

We note that ^{120}Sn nucleus contains 50 protons, which is a magic nucleon number. Therefore, ^{120}Sn nucleus,

which requires 10.6 MeV for removing a proton from it,
is more stable than ^{121}Sb nucleus, which requires 5.8
MeV for removing a proton from it.

